

Chemistry 1094 - Chemical and Physical Changes

Date: _____ **Name:** _____ **Station #:** _____

OBJECTIVE: To recognize when a chemical reaction has occurred and to make deductions based upon careful observations.

PROCEDURE: As in the Chemistry 1094 Lab Manual, pages 53.

OBSERVATIONS/DATA: see attached sheets

CONCLUSION: What evidence can be used to determine whether or not a chemical reaction has occurred? Give examples from the reactions performed today

Test	Solutions	Observations	Deductions (choose one from each column below)	
1	Calcium chloride and Sodium carbonate		Chemical change Physical change Neither	Reaction No reaction
2	Silver nitrate and Copper		Chemical change Physical change Neither	Reaction No reaction
3	6 M Sodium hydroxide and 6 M Hydrochloric acid		Chemical change Physical change Neither	Reaction No reaction
4	Hydrogen peroxide and Potassium iodide		Chemical change Physical change Neither	Reaction No reaction

Test	Solutions	Observations	Deductions (choose one from each column below)	
5	Ferric chloride and Ammonium thiocyanate		Chemical change Physical change Neither	Reaction No reaction
6	Step 5 and Silver nitrate		Chemical change Physical change Neither	Reaction No reaction
7	Potassium permanganate and water		Chemical change Physical change Neither	Reaction No reaction
8	Potassium permanganate and oxalic acid		Chemical change Physical change Neither	Reaction No reaction
9	Potassium permanganate and oxalic acid and heat		Chemical change Physical change Neither	Reaction No reaction

Test	Solutions	Observations	Deductions (choose one from each column below)	
10	Sucrose and Heat		Chemical change Physical change Neither	Reaction No reaction
11	Sodium Chloride and Heat		Chemical change Physical change Neither	Reaction No reaction
12	Tin and Heat		Chemical change Physical change Neither	Reaction No reaction
13	Ice and Heat		Chemical change Physical change Neither	Reaction No reaction

Remember to write a conclusion (see first page of handout).