

CHEMICAL REACTIONS

Date: _____ Name: _____ Lab Day/Time: _____

Objectives

- to make careful observations of chemical reactions using the senses of sight, touch, and smell,
- to describe the chemical reactions that have taken place in terms of balanced chemical equations. (Find out from your lab instructor whether you should do this second objective.), and
- to identify unknown ions by using precipitation reactions.

Procedure

As in Chem 1094 lab manual, pp. _____

Observations and Chemical Reactions

Part A: Chemical Reactions

PROCEDURE 1

To 5 drops of lead (II) nitrate solution add 5 drops of potassium iodide solution.

Observations

Equation



PROCEDURE 2

To 5 drops of lead (II) nitrate solution add 15 drops of hydrochloric acid solution.

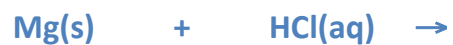
Observations

Equation



PROCEDURE 3

To a 1 cm strip of magnesium metal add 3mL of hydrochloric acid solution

Observations**Equation****PROCEDURE 4**

To 3mL of sodium hydroxide solution add 3 mL of hydrochloric acid solution

Observations**Equation****PROCEDURE 5**

To a piece of zinc metal add 3mL of copper (II) sulphate solution.

Observations**Equation**

PROCEDURE 6

To 10 drops of sodium chloride solution add 5 drops of silver nitrate solution.

Observations**Equation****PROCEDURE 7**

To 5 drops of barium chloride solution add 5 drops of sodium carbonate solution.

Observations**Equation****PROCEDURE 8**

To 5 drops of barium nitrate solution add 5 drops of sodium sulphate solution.

Observations**Equation**

PROCEDURE 9

To 10 drops of sodium carbonate solution add 1 mL of hydrochloric acid solution.

Observations**Equation****PROCEDURE 10**

To 10 drops of sodium carbonate solution add 5 drops of copper (II) sulphate solution.

Observations**Equation**

Part B: Physical Processes

PROCEDURE 11

To 3mL of tap water, add 1 gram of solid ammonium chloride.

Observations

Equation

PROCEDURE 12

To 3mL of water, add 1 gram of solid calcium chloride.

Observations**Equation****PROCEDURE 13**

To 5 drops of nickel chloride solution add 10 drops of sodium chloride solution.

Observations**Equation**

Chemical Process:

PROCEDURE 14

Dip a clean piece of copper wire into a thin layer of silver nitrate solution which is sitting in a watchglass under a stereomicroscope. Observe the results through the stereomicroscope.

Observations

Equation

Questions

Attach any assigned questions